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2 (f) At temperatures below the value calculated in part 2 (e), decomposition of NO into its elements should be spontaneous. However, in car exhausts this decomposition reaction does **not** take place in the absence of a catalyst. Suggest why this spontaneous decomposition does **not** take place.

A High Activation Energy

(1 mark)

2 (g) A student had an idea to earn money by carrying out the following reaction.

$$\text{C}(\text{graphite}) \rightarrow \text{C}(\text{diamond})$$

Use data from the table on page 4 to calculate values for ΔH° and ΔS° for this reaction. Use these values to explain why this reaction is **not** feasible under standard pressure at any temperature.

$\Delta G > 0$
 $\Delta G = \Delta H - T\Delta S$

ΔH° _____
 ΔS° _____

Explanation _____

(3 marks)

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