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so many fake sites. this is the first one which worked! Many thanks

**Strong versus Weak Acids**  
The following are strong acids:

**What**  
The following are strong acids: hydrochloric acid, nitric acid, sulfuric acid, perchloric acid, hydrobromic acid, and hydroiodic acid. These acids are completely ionized in water. The following are weak acids: acetic acid, carbonic acid, phosphoric acid, and many others. These acids are only partially ionized in water.

**Model 1 - Acid Strength and Conductivity**

The diagram shows four beakers. The first beaker is labeled 'Strong' and contains many small circles representing ions. The second beaker is labeled 'Weak' and contains many small squares representing ions. The third beaker is labeled 'Weak' and contains many small squares representing ions. The fourth beaker is labeled 'Strong' and contains many small circles representing ions. A legend indicates that a square represents a weak acid and a circle represents a strong acid.

1. Explain the conductivity differences shown above.

2. Explain the conductivity differences shown above.

3. Explain the conductivity differences shown above.

4. Explain the conductivity differences shown above.

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**Answer Key Pogil Strong Versus Weak Acids**